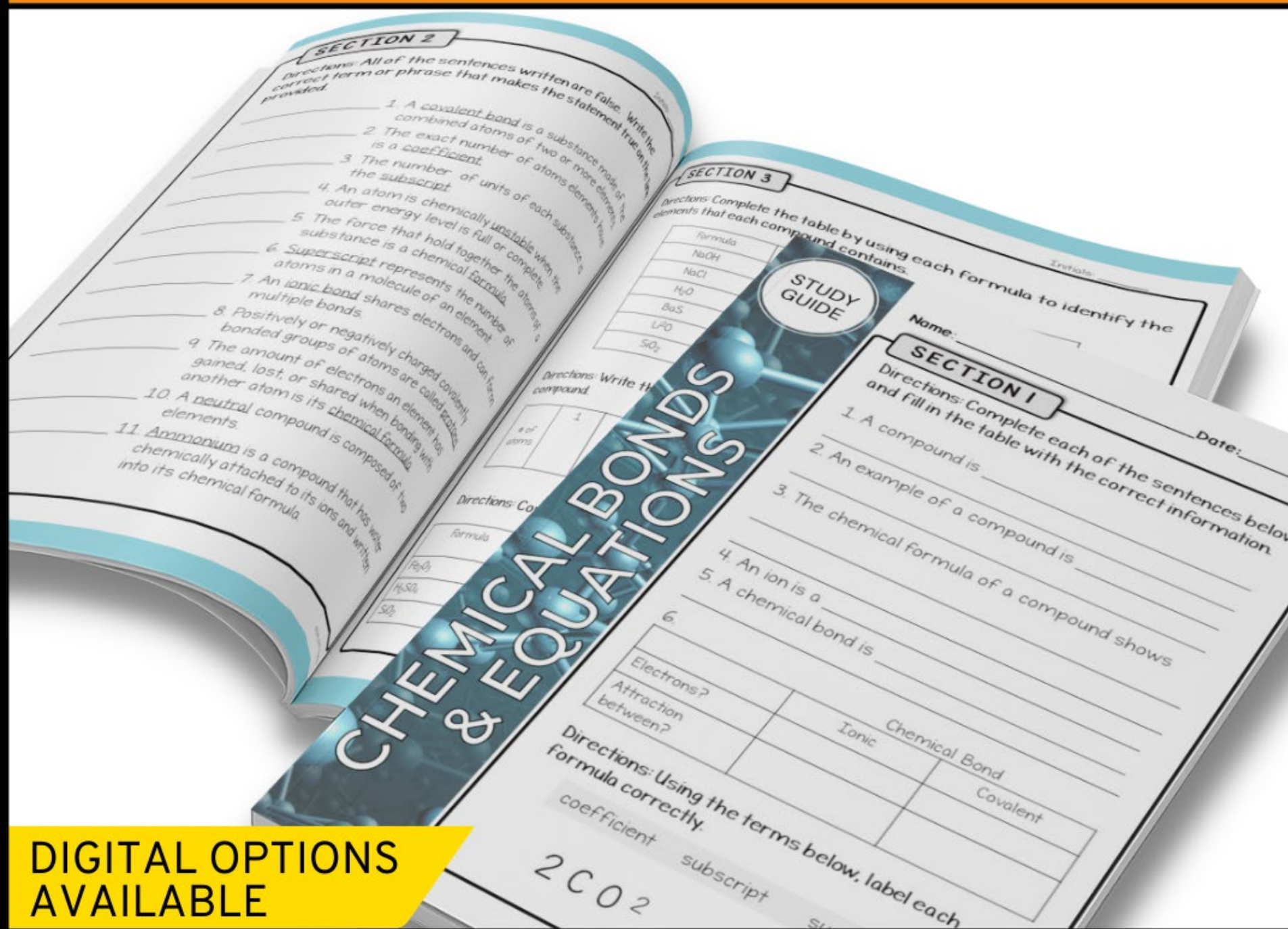




DIGITAL OPTIONS
AVAILABLE

STUDY GUIDE

REVIEW STRATEGIES · CRITICAL THINKING

Includes **PRINT** or **DIGITAL** Options

Date: _____

SECTION 1

Directions: Complete each of the sentences below and fill in the table with the correct information.

1. A compound is _____
2. An example of a compound is _____
3. The chemical formula of a compound shows _____
4. An ion is a _____
5. A chemical bond is _____
6. _____

	Chemical Bond
	Ionic
	Covalent

Initials: _____

SECTION 2

Directions: All of the sentences written are false. Write the correct term or phrase that makes the statement provided.

1. A covalent bond is a substance.
2. The exact number of atoms is a coefficient.
3. The number of units of each is the subscript.
4. An atom is chemically unstable.
5. The force that hold together outer energy level is full or covalent.
6. Superscript represents the number of atoms in a molecule of an element.
7. An ionic bond shares electrons to form multiple bonds.

Initials: _____

SECTION 3

Directions: Complete the table by using each formula to identify the elements that each compound contains.

Formula	Element 1	Element 2
NaOH		
NaCl		
H ₂ O		
BaS		
Li ₂ O		
SiO ₂		

Directions: Write the prefixes for each number of binary covalent compound.

	1	2	3	4	5	6	7	8

Initials: _____

SECTION 4

Directions: Fill in the graphic organizer below with the correct definition for each term.

Chemical Reactions

Reactants

Law of conservation of mass states _____

Directions: Complete the table by writing or drawing information in the missing boxes.

Symbol	Meaning
→	Electric current
Light →	Reactants heat
(aq)	Solid
+	
(l)	

Initials: _____

SECTION 5

Directions: Answer the questions below using this equation.

$$\text{NiCl}_2(aq) + 2\text{NaOH}(aq) \rightarrow \text{Ni}(\text{OH})_2(s)$$

1. What is a chemical equation?

2. What are the reactants in this equation?

3. What are the products?

4. What do the coefficients tell you about the reactants?

5. What is the physical state of the products?

6. Write the equation above as a sentence.

Initials: _____

SECTION 6

Directions: Answer the questions below and then fill in the graphic organizer with the correct steps used in balancing an equation.

How do you know when an equation is balanced?

What should you never do when balancing equations?

Step 1

Step 2

Step 3

Step 4

Directions: Balance this equation.

$$\text{Li}(s) + \text{H}_2\text{O} \rightarrow \text{LiOH}(aq) + \text{H}_2$$



NGS Magnified Study Guides are directly aligned to the notes and assessments offered by NGS Magnified and include a variety of review strategies that meet the needs of your learners for independent study and indirect instruction.

Each study guide provides a combination of strategies which may include:

graphic organizers

critical thinking

vocabulary building

theme connection

compare and contrast

matching

problem solving

fill-in-the-blank

concept mapping

short answer

interpreting data

real world application

QR videos with accompanying questions



Thank you for sharing NGS Magnified with your students!

Terms of Use

Copyright © NGS Magnified, LLC (formerly Nitty Gritty Science, LLC.) All rights reserved by author Dr. Erica Colón. This product is to be used by the original downloader only. Copying for more than one teacher, classroom, department, school, or school system is prohibited. This product may not be distributed or displayed digitally for public view. Failure to comply is a copyright infringement and a violation of the Digital Millennium Copyright Act (DMCA). Clipart and elements found in this PDF are copyrighted and cannot be extracted and used outside of this file without permission or license. Intended for classroom and personal use ONLY.

Contact Information:

Email: admin@nittygrittyscience.com

Website: www.NGSmagnified.com

TPT: <https://www.teacherspayteachers.com/Store/Nitty-Gritty-Science>

